

Enhancing the in vivo transdermal delivery of gold nanoparticles using poly(ethylene glycol) and its oleylamine conjugate

Pa Fan Hsiao¹⁻³
Sydney Peng⁴
Ting-Cheng Tang⁴
Shuan-Yin Lin⁵
Hsieh-Chih Tsai⁴

¹Department of Dermatology, Mackay Memorial Hospital, ²Mackay Medicine, Nursing and Management College, ³Mackay Medical College, New Taipei City, ⁴Graduate Institute of Applied Science and Technology, National Taiwan University of Science and Technology, Taipei, ⁵National Applied Research Laboratories, Instrument Technology Research Center, Hsinchu, Taiwan

Abstract: In this study, we investigated the effect of (ethylene glycol) (PEG) and PEG–oleylamine (OAm) functionalization on the skin permeation property of gold nanoparticles (GNS) in vivo. Chemisorption of polymers onto GNS was verified by a red shift in the ultraviolet–visible spectrum as well as by a change in the nanoparticle surface charge. The physicochemical properties of pristine and functionalized nanoparticles were analyzed by ultraviolet–visible spectroscopy, zeta potential analyzer, and transmission electron microscopy. Transmission electron microscopy revealed that the interparticle distance between nanoparticles increased after GNS functionalization. Comparing the skin permeation profile of pristine and functionalized GNS, the follicular deposition of GNS increased twofold after PEG–OAm functionalization. Moreover, PEG- and PEG–OAm-functionalized nanoparticles were able to overcome the skin barrier and deposit in the deeper subcutaneous adipose tissue. These findings demonstrate the potential of PEG- and PEG–OAm-functionalized GNS in serving a multitude of applications in transdermal pharmaceuticals.

Keywords: skin penetration, amphiphilic copolymer, gold nanoparticle, oleylamine, poly(ethylene glycol)

Introduction

The use of nanotechnology in the form of nanomedicine for prevention, diagnosis, and treatment of diseases, including cancer, is one branch of this fast developing field.¹ Breakthroughs in particle size control and nanoparticle modification have created the possibility of designing nanoparticles that hold enormous potentials in drug delivery. Increased therapeutic efficacy is often achieved by improving selectivity toward target cells and preventing toxic effects on normal cells.² However, coupled with advancements in nanomedicine are critiques on the safety thereto. With direct systemic introduction, systemic and local toxicities are of tremendous concern as nanoparticles may aggregate and cause subsequent tissue inflammation.³ Gold nanoparticles (GNS) in the size range of 18–37 nm caused major damage to the liver, spleen, and lungs,⁴ while PEG-functionalized GNS accumulated in the liver after injection and inflicted both acute inflammation and cellular damage.⁵ On the contrary, transdermal delivery exhibits relatively lower risks of systemic toxicity due to its controlled depot effect, which permits the use of relatively potent drugs.^{6,7} Furthermore, transdermal delivery is particularly attractive because of its large surface area, ease of administration, noninvasiveness, and ability to bypass the first-pass effect.⁸

Correspondence: Hsieh-Chih Tsai
Graduate Institute of Applied Science and Technology, National Taiwan University of Science and Technology, No 43, Keelung Road, Section 4, Daan District, Taipei 10607, Taiwan
Tel +886 2 2730 3625
Fax +886 2 2730 3733
Email h.c.tsai@mail.ntust.edu.tw



In transdermal delivery, drug permeability is challenged by the many layers of the skin. The stratum corneum (SC), a 10–20 nm skin layer composed of tightly packed sheets of metabolically inactive cells surrounded by a lipid extracellular matrix, acts as the foremost barrier against transdermal delivery.^{9,10} Physicochemical properties, such as size, surface charge, and material composition, are factors that dictate the permeation ability of nanoparticles.^{11,12} Because cells are so densely packed inside the hydrophobic SC, passage of charged, hydrophilic, and large drug molecules is extremely limited.¹³ Nevertheless, hydrophilic pharmaceutical compounds can enter the skin through hair follicles and tight junctions, which are skin barrier regulators in the viable epidermis. The follicular delivery route has been an attractive target for particulate drug delivery systems as they increase drug permeation and act as depot for sustained therapeutics.^{14,15} On the other hand, multiple physical and chemical skin enhancers have been explored, ranging from microneedle technology, which deliberately creates microsized holes in the SC,^{16–18} to the addition of amines and fatty acids to modify SC properties so they become more prone to permeation.^{19,20} More specifically, enhancing SC permeability by targeting specific SC lipids has long been the strategy used in the design of transdermal systems.^{21,22}

Silver- and gold-based metal nanoparticles have been shown in the literature to permeate the skin. However, due to different experimental parameters and the limited number of studies, a definite conclusion regarding their permeation ability cannot be drawn with simplicity. By using a Franz cell setup, one group reported that polyvinylpyrrolidone-coated silver nanoparticles entered intact skin while permeating damaged skin *in vitro*.²³ GNS, on the

other hand, have gained considerable attention because of their simple preparation, controllable size and distribution, biocompatibility, and functionalizing potential.²⁴ Furthermore, GNS demonstrate distinct electrical and chemical properties, such as high electron density, which make them suitable for advanced applications, such as biosensing, diagnostic imaging, and photodynamic therapy.^{25–27} A recent study fully evaluated the *in vitro* skin permeation of pegylated and cell-penetrating peptide-modified GNS through different skin layers using transmission electron microscopy (TEM).²⁸

Despite numerous discussions on the *in vitro* skin permeation of GNS, those focusing on their *in vivo* permeation are relatively limited.^{28–31} Studies of GNS as contrasting agents after topical application are available for several animal models.^{32,33} However, the emphasis of these studies is on increasing the optical intensity rather than drawing conclusions on their unassisted permeation properties. This study sets out to investigate the *in vivo* permeation of GNS functionalized with an amphiphilic permeation enhancer. Here, we used 10 nm GNS as the starting material and amphiphilic poly(ethylene glycol) (PEG)–oleylamine (OAm) as the chemical enhancer. The amphiphilic PEG–OAm was prepared by reaction between the carboxyl group of PEG and the terminal amine group of OAm, as shown in Figure 1. The physical, chemical, and permeation properties of pristine, PEG-functionalized, and PEG–OAm-functionalized derivatives were investigated. Localization of nanoparticles was observed by fluorescent microscopy and TEM. This study points to provide insights into understanding the effects of PEG and OAm on skin permeation *in vivo* and presents the system as a potential transdermal carrier.

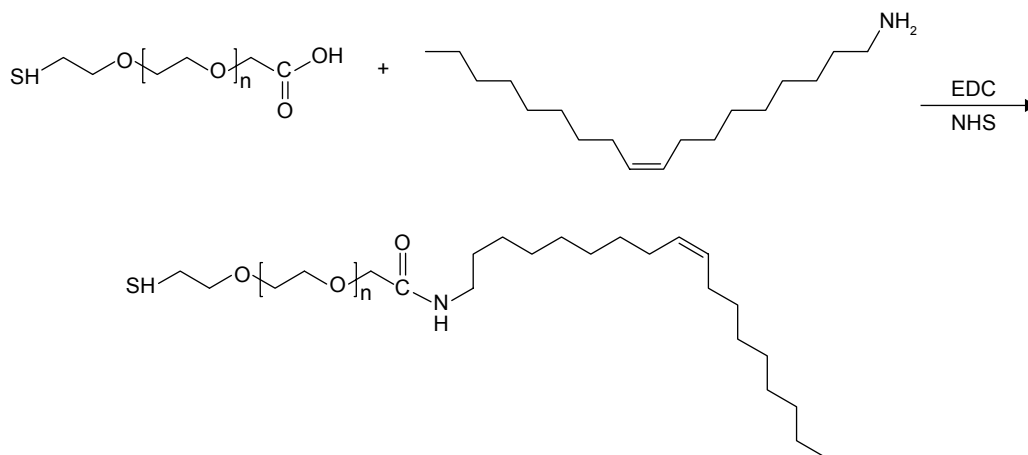


Figure 1 Synthesis of amphiphilic copolymer (SH-PEG-OAm).

Abbreviations: EDC, 1-ethyl-3-(3-dimethylaminopropyl)-carbodiimide; NHS, N-hydroxysuccinimide; OAm, oleylamine; PEG, poly(ethylene glycol).

Materials and methods

Materials

Citric ion-stabilized GNS (10 nm), PEG 2-mercaptoethyl ether acetic acid (HS-PEG-COOH) (MW ~2,100), OAm, *N*-hydroxysuccinimide (NHS), 1-ethyl-3-(3-dimethylaminopropyl)-carbodiimide (EDC), glutaraldehyde solution, fluorescein isothiocyanate (FITC), osmonium tetroxide solution, uranyl acetate, and lead citrate were obtained from Sigma-Aldrich Co. (St Louis, MO, USA). Paraformaldehyde and epoxy resin embedding medium were purchased from Acros (Geel, Belgium) and Electron Microscopy Sciences (Hatfield, PA, USA), respectively. All other chemicals were used as received unless otherwise specified. Ethical approval was not required or sought from the IRB of Mackay Memorial Hospital, Taiwan because the experiments are not performed on humans.

Coupling of HS-PEG-COOH with OAm

PEG-OAm was prepared by EDC/NHS coupling chemistry where the PEG carboxylic acid was activated with EDC and subsequently reacted with the primary amine on OAm to afford an amide bond. Briefly, HS-PEG-COOH, OAm, EDC, and NHS were reacted overnight in phosphate-buffered solution at a molar ratio of 1:1:1:1. Upon completion of reaction, excess EDC/NHS and residual OAm were removed by dialysis (molecular cutoff [MWCO]: 1,000 Da).

PEG-OAm coupling was verified by ¹H nuclear magnetic spectroscopy (NMR) and Fourier transformed infrared spectroscopy (FT-IR). Unreacted OAm, PEG, and PEG-OAm were dissolved in dimethyl sulfoxide-*d*₆ and analyzed by ¹H NMR using an AVANCE III HD 600 MHz spectrometer (Bruker Corporation, Billerica, MA, USA). FT-IR samples were prepared on CaF₂ disks using the drop casting method and analyzed on a Nicolet 6700 spectrometer (Thermo Fisher Scientific, Waltham, MA, USA).

Surface modification of GNS

PEG functionalization of GNS was performed by direct mixing of GNS dispersion with PEG aqueous solution. In brief, a pristine GNS dispersion at 1 mM (4 mL) was added to a 1 mM HS-PEG-COOH aqueous solution (1 mL) under constant stirring and reacted overnight at 25°C. Excess PEG was removed by dialysis against deionized water (MWCO: 3,500 Da) for 24 hours to afford GNS-PEG. A similar procedure was used to synthesize PEG-OAm-functionalized GNS. In short, a pristine GNS dispersion (1 mM) was added to a 4 mM of PEG-OAm aqueous solution and stirred overnight at 25°C to provide PEG-OAm-functionalized GNS (GNS-PEG-OAm) in aqueous solution. Unreacted PEG-OAm was removed by dialysis (MWCO: 3,500 Da).

Verification of PEG and PEG-OAm functionalization of GNS

Transparent dispersions of pristine GNS, GNS-PEG, and GNS-PEG-OAm were deposited on CaF₂ disks by the drop casting method and air-dried for 12 hours prior to FT-IR analysis.

Particle size, particle morphology, and zeta potential

The morphology and dispersion of pristine and functionalized GNS were visualized using TEM (Hitachi H-7000; Hitachi Ltd., Tokyo, Japan). For sample preparation, 200 mM nanoparticle dispersions (5 mL) were drop casted onto copper grids (300 mesh) and air-dried. The hydrodynamic particle size, polydispersity index, and surface charge of nanoparticles were recorded using a Horiba SZ-100 system (HORIBA Ltd., Kyoto, Japan). Each zeta potential value (mV) represents an average of three runs of ten measurements each.

Optical property analysis

Optical property of the nanoparticles was studied by ultraviolet-visible (UV-Vis) spectroscopy. After centrifugation of GNS suspensions, collected nanoparticles were redispersed in deionized water and transferred to quartz cuvettes. Absorbance spectra were recorded between 200 nm and 1,000 nm using a V-670 spectrophotometer (Jasco, Tokyo, Japan).

Preparation and detection of FITC-labeled GNS

In order to visualize the location of GNS after topical application, GNS were labeled with FITC. Briefly, FITC solution at pH 9 was reacted with the various GNS dispersions for 12 hours at 4°C. The products were dialyzed against deionized water (MWCO: 3,500 Da) for 24 hours to remove unreacted FITC. Fluorescent intensity was confirmed using an F-7000 spectrometer (Hitachi Ltd.).

In vivo permeation of GNS

Animal studies were conducted at Mackay Memorial Hospital (Taipei, Taiwan) using three Sprague Dawley rats weighing 5,000–6,000 g. For animal experiments, Institutional Animal Care and Use Committee (IACUC) guidelines were followed. Anesthesia was induced by intraperitoneal injection of Zoletil (40 mg/kg). Dorsal hair was clipped, and dressings presoaked with 20 μL FITC-labeled GNS (30 nM) were topically applied onto test areas (2×2 cm) for 24 hours. Three different formulations, pristine GNS, GNS-PEG, and GNS-PEG-OAm, were

applied to each animal. Test animals were sacrificed by CO₂ asphyxiation, and the test areas were washed with phosphate-buffered saline (PBS) prior to incision. The collected skin tissues were washed with PBS again and fixed with 4% paraformaldehyde (0.1 M PBS, pH 7.4) for 24 hours at 4°C.

Tissue samples were removed from paraformaldehyde and dehydrated in 30% sucrose until they sank to the bottom of the vial. Next, the samples were washed three times with PBS and stored in embedding cassettes overnight at -20°C before they were embedded in O.C.T (Tissue-Tek) and frozen at -18°C for 20 minutes. Frozen samples were sectioned perpendicular to the knife from the hypodermis up through the epidermal layers to prevent contamination from surface to deeper regions. The thickness of confocal laser scanning microscopy (CLSM) samples was ~30–60 μm. Sections were mounted onto glass slides, and the relative permeation profile was observed by CLSM (TCS SP2; Leica Microsystems, Wetzlar, Germany). A series of five random CSLM images were taken, analyzed, and averaged. Leica Application Suite X provided by Leica Microsystems was used to extract the z-axis fluorescence intensity profile from the selected region of interest.

Samples for TEM analysis were stained with OsO₄ and dehydrated prior to infiltration using epoxy resin. Samples were cured, and 100 nm thick sections were cut using an ultramicrotome (EM UC7; Leica Microsystems), mounted

on copper TEM grids, and stained with uranyl acetate and lead citrate for 20 minutes and 5 minutes, respectively. The deposition of nanoparticles in the subcutaneous adipose tissue was visualized.

Results and discussion

Coupling of SH-PEG-COOH with OAm

Figure 2 shows the ¹H NMR spectrum of PEG–OAm with ten characteristic signals observed. In particular, a triplet was detected between 0.83 ppm and 0.86 ppm (peak a), corresponding to the methyl group of OAm. Signals at 1.23–1.33 ppm (t, peaks b), 1.93–1.99 ppm (m, peak c), and 2.71 ppm (t, peak e) are attributed to OAm methylenic protons. One multiplet was recorded at 5.31–5.36 ppm, which is consistent with the methinic protons of OAm (m, peak E). In addition, signals at 2.28–2.29 ppm (m, peak 1), 2.61–2.64 ppm (m, peak 2), 3.5 ppm (m, peak 3), 3.89–3.91 ppm (t, peak 4), and 3.9 ppm (s, peak 5) may be assigned to PEG methylenic protons.

Figure 3 represents the FT-IR spectra of the starting materials and PEG–OAm. Distinct absorption peaks corresponding to terminal functional groups are visible at 3,300–3,500 cm⁻¹ (–OH) for PEG as well as at 3,400 cm⁻¹ and 3,250 cm⁻¹ (–NH₂) for OAm prior to reaction. Carbonyl absorption at 1,700 cm⁻¹ confirms the successful coupling of PEG and OAm through amide bond formation. Emergence

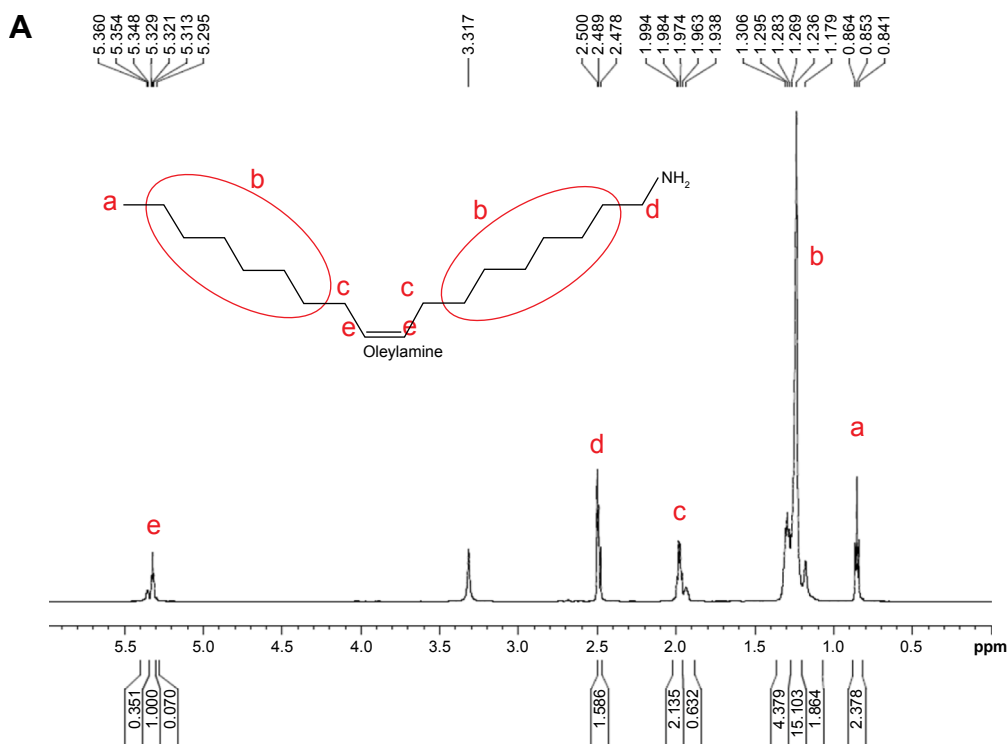


Figure 2 (Continued)

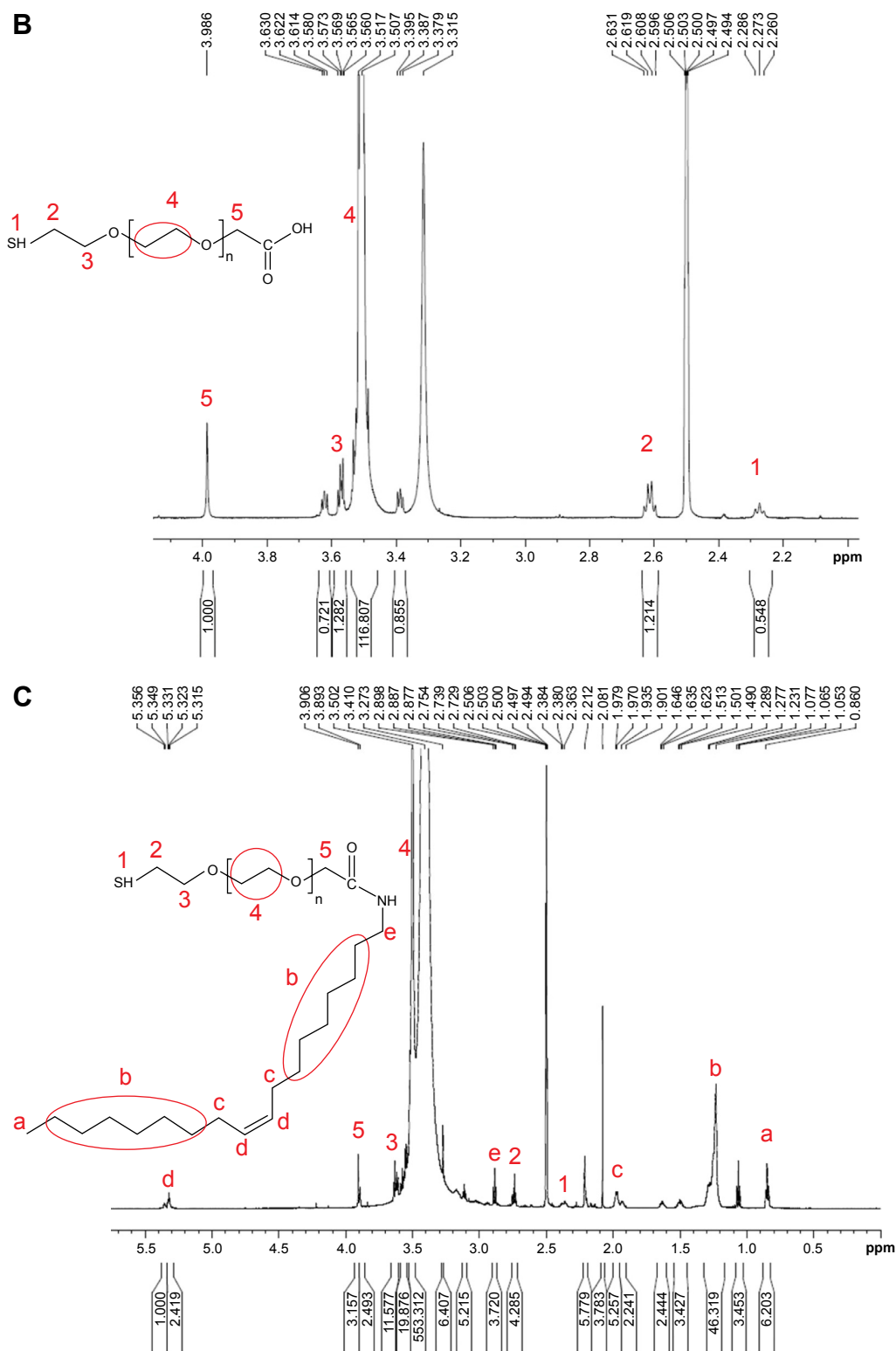


Figure 2 ^1H NMR spectrum of oleylamine (A), SH-PEG-COOH (B), and coupled SH-PEG-OAm (C).

Note: The letters a–e represent the protons in oleylamine; the numbers 1–5 represent the protons in SH-PEG-COOH.

Abbreviations: NMR, nuclear magnetic spectroscopy; OAm, oleylamine; PEG, poly(ethylene glycol).

of the NH bending vibration at $1,590\text{ cm}^{-1}$ and disappearance of $-\text{OH}$ and $-\text{NH}_2$ vibrations further validate PEG-OAm coupling. These observations indicate that PEG-OAm composed of both PEG and OAm domains.

Characterization of pristine and functionalized GNS

FT-IR spectra of pristine GNS, GNS-PEG, and GNS-PEG-OAm are shown in Figure 4. For pristine GNS, characteristic

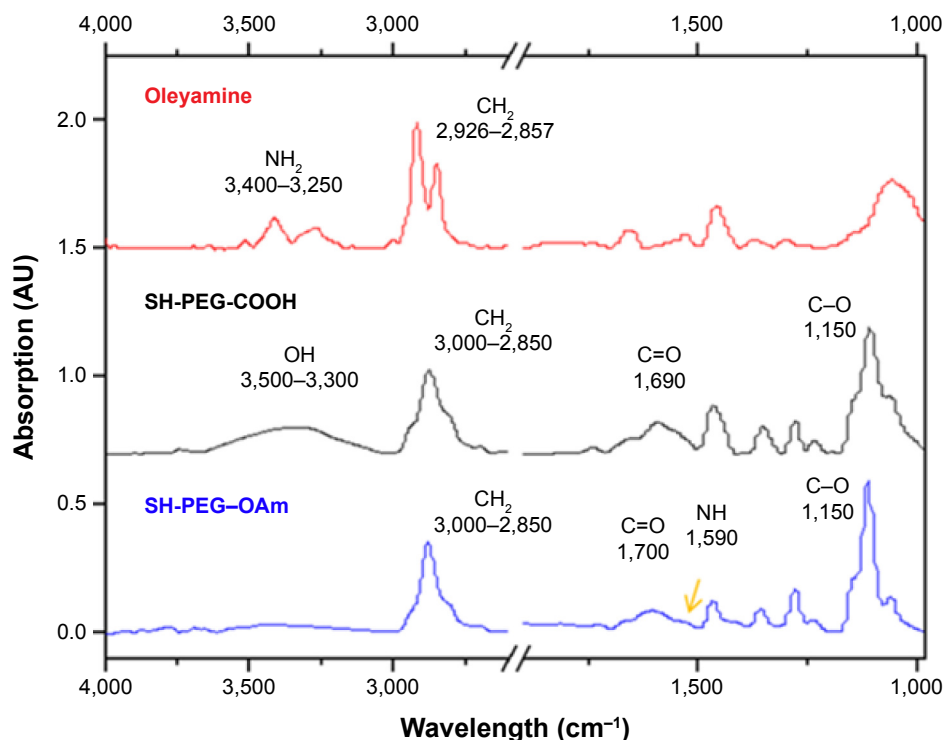


Figure 3 FT-IR spectra of oleylamine, SH-PEG-COOH, and coupled SH-PEG-OAm.

Note: Coupling is verified by the appearance of NH bending vibration at $1,590\text{ cm}^{-1}$ and loss of terminal $-\text{OH}$ and $-\text{NH}_2$ absorptions. The yellow arrow indicates the formation of amide bond in SH-PEG-OAm.

Abbreviations: FT-IR, Fourier transformed infrared spectroscopy; OAm, oleylamine; PEG, poly(ethylene glycol).

absorption bands at $1,700\text{--}1,750\text{ cm}^{-1}$ and $3,450\text{ cm}^{-1}$ correspond to $-\text{COO}^-$, $\text{C}=\text{O}$, and $-\text{OH}$ groups, respectively, in the surface-stabilizing citric acids. Reaction with SH-PEG-COOH resulted in the appearance of $\text{C}=\text{O}$ and $\text{C}-\text{O}$ stretching

bands at $1,690\text{ cm}^{-1}$ and $1,150\text{ cm}^{-1}$, respectively. The strong absorption peak at $2,850\text{ cm}^{-1}$ represents CH_2 stretching frequency of PEG and verifies the successful functionalization of PEG onto GNS surface. Furthermore, the disappearance

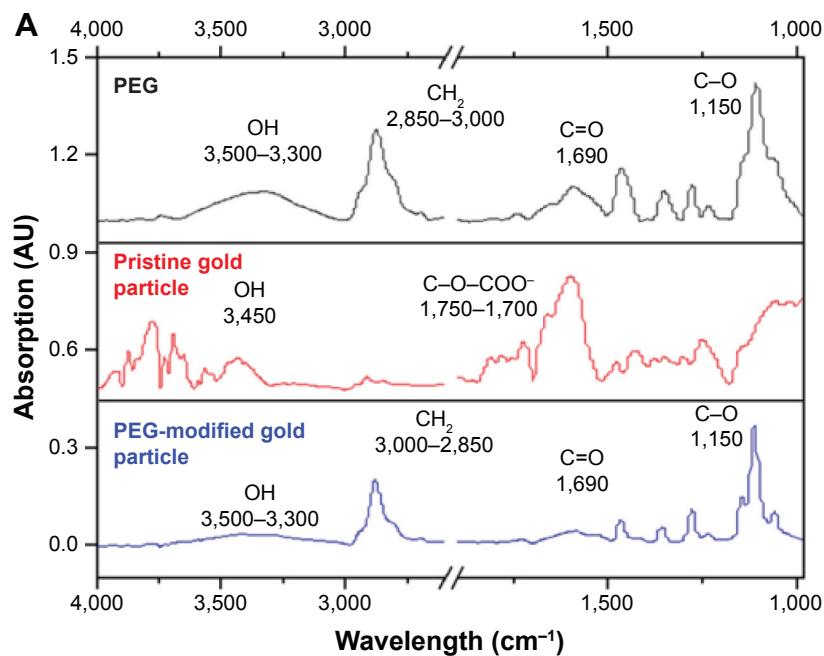


Figure 4 (Continued)

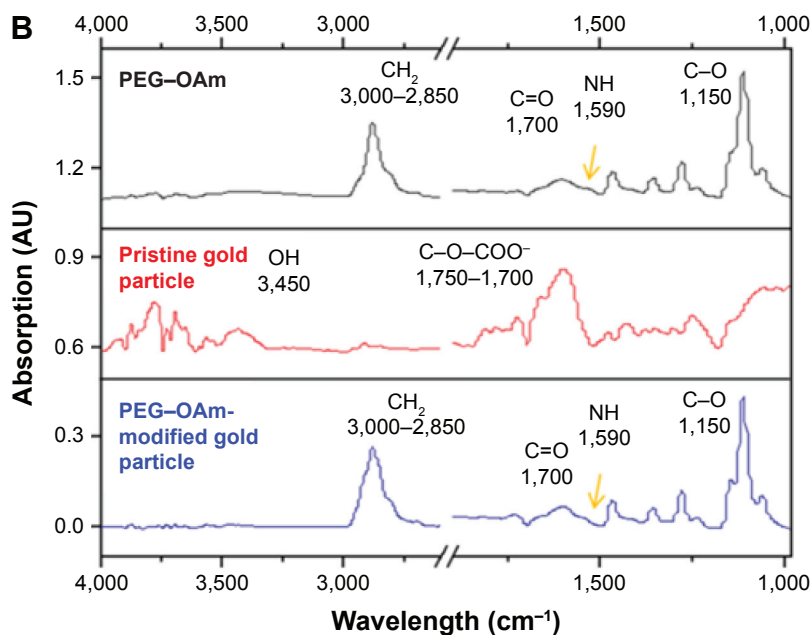


Figure 4 FT-IR spectra of (A) PEG-modified GNS and its starting materials and (B) PEG-OAm-modified GNS and its starting materials.

Notes: Appearance of strong CH_2 stretching frequency at $3,000\text{--}2,850\text{ cm}^{-1}$ polymer functionalization. The yellow arrow indicates the amide bond from PEG-OAm.

Abbreviations: FT-IR, Fourier transformed infrared spectroscopy; GNS, gold nanoparticles; OAm, oleylamine; PEG, poly(ethylene glycol).

of the --OH stretching band at $3,300\text{--}3,500\text{ cm}^{-1}$ of carboxyl group of PEG suggests the formation of PEG-functionalized GNS. This functionalization process was driven by the high affinity of thiol groups toward gold surfaces, forming Au-sulfur bonds. Similarly, PEG-OAm-functionalized GNS exhibited characteristic absorption bands of PEG-OAm at $1,690\text{ cm}^{-1}$ (C=O) and $1,150\text{ cm}^{-1}$ (C-O). The --NH bending vibration at $1,590\text{ cm}^{-1}$ confirms PEG-OAm conjugation onto GNS.

Particle morphology, zeta potential, hydrodynamic size, and UV-Vis absorption

As GNS have relatively high electron density, they are excellent probes for electron microscopy, and their morphology can be assessed directly by TEM, as shown in Figure 5. All nanoparticles were spherical in shape with diameters of $\sim 10\text{ nm}$. The size of GNS, as observed under TEM, did not increase after functionalization since only the gold core is visible at the acceleration voltage used.³⁴ As shown in Figure 5A and B, interparticle distance increased with PEG functionalization, which provided a protective layer around the nanoparticle to allow free dispersion in water (Figure 5B).³⁵ Incorporation of hydrophobic OAm did not compromise dispersity and provided highly monodispersed nanoparticles, thereby increasing the number of free nanoparticles available for interacting with biological membranes (Figure 5C).³⁶

Since the PEG-OAm is an amphiphilic copolymer, it still can disperse well in water. The OAm self-assembled and attached to the hydrophobic surface of GNS and then simultaneously possessed PEG segment located at the exterior part. A similar system was also reported in our previous study.³⁷ The dispersion morphology of PEG-OAm functionalization of GNS is shown in Figure 5D.

As shown in Figure 6A, pristine GNS are negatively charged due to surface citrates as stabilizers. After PEG functionalization, GNS zeta potential increased from $-70.21 \pm 2.0\text{ mV}$ to $-28.53 \pm 1.4\text{ mV}$, a result of PEG shielding the negatively charged citrate ions. Functionalization with PEG-OAm further increased the surface charge to $+10.1 \pm 1.7\text{ mV}$. These results confirm the stable incorporation of PEG and PEG-OAm by formation of thiol-Au bonds.

The dynamic light scattering size and polydispersity index were also confirmed by dynamic light scattering (Table 1). The results indicated that PEG-modified GNS exhibited a large hydrodynamic diameter. For PEG-OAm GNS, the hydrodynamic size slightly reduced when compared with PEG-modified GNS. It is because that the hydrophobic interaction between the gold surface and OAm segment results in the decrease in hydrodynamic particle size. This result is also consistent with the expected morphology of PEG-OAm GNS in water. In addition, pristine GNS, PEG-GNS, and PEG-OAm GNS exhibited good polydispersity index. Since absorption bands are directly related to the aspect ratio of

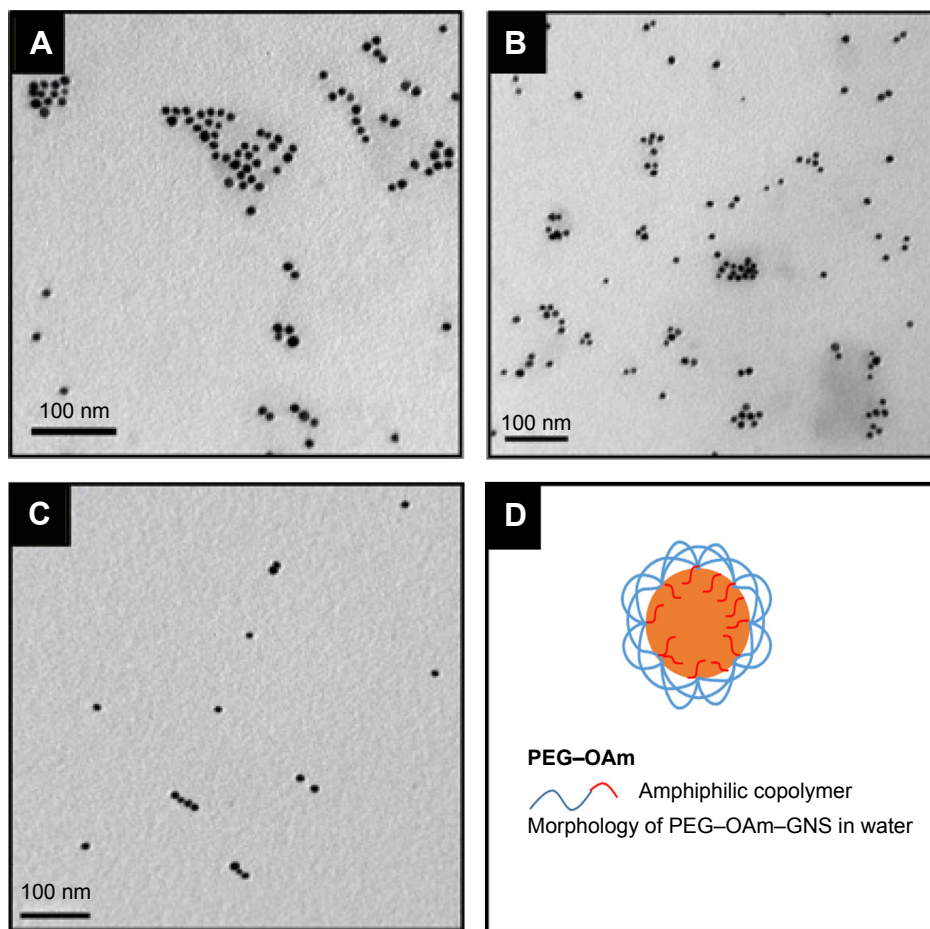


Figure 5 Morphology and dispersion of gold nanoparticles in this study.
Notes: TEM images of (A) pristine GNS, (B) GNS-PEG, and (C) GNS-PEG-OAm. Increase in interparticle distance is observed after functionalization. (D) The dispersion of amphiphilic copolymer, PEG-OAm, and functionalization of GNS in water. Magnification: 100,000× for (A), (B) and (C).
Abbreviations: GNS, gold nanoparticles; OAm, oleylamine; PEG, poly(ethylene glycol); TEM, transmission electron microscopy.

metal nanoparticles, UV-Vis is most often used for GNS characterization. As shown in Figure 6B, a red shift of surface plasmon resonance (SPR) occurred after PEG and PEG-OAm functionalization, which indicates particle size increase

due to formation of a protective layer on the surface of GNS. This is concurrent with findings in TEM. Size-related SPR shift due to changes in electromagnetic interactions between GNS and incident of light is described by the Mie theory.³⁸

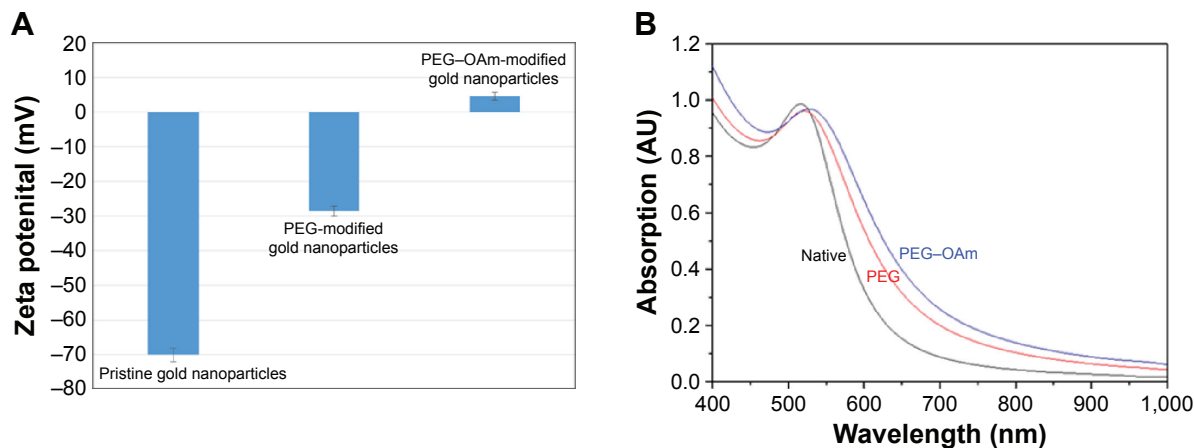


Figure 6 (A) Zeta potential and (B) UV-Vis spectra of pristine GNS, GNS-PEG, and GNS-PEG-OAm.
Notes: Error bars represent SD, n=3. A red shift in the spectrum occurred after functionalization indicating growth in particle size.
Abbreviations: GNS, gold nanoparticles; OAm, oleylamine; PEG, poly(ethylene glycol); UV-Vis, ultraviolet-visible.

Table I The hydrodynamic diameter of pristine gold nanoparticle, PEG-modified gold nanoparticle, and PEG–OAm-modified gold nanoparticle

| Types of gold nanoparticle | Pristine gold nanoparticle | PEG-modified gold nanoparticle | PEG–OAm-modified gold nanoparticle |
|----------------------------|----------------------------|--------------------------------|------------------------------------|
| DLS size (nm) | 25±2 | 58±4 | 41±3 |
| PDI | 0.105 | 0.152 | 0.113 |

Abbreviations: DLS, dynamic light scattering; OAm, oleylamine; PDI, polydispersity index; PEG, poly (ethylene glycol).

The bulky PEG layer on GNS elicited a 2 nm red shift in the main localized SPR. Furthermore, the narrow plasmon band at 532 nm and the lack of absorbance peak >600 nm also indicate uniform size distribution and stable dispersion without particle aggregation, respectively.

In vivo skin permeation assessment

Confocal microscopy was used to visualize fluorescein distribution in the skin layers after pristine GNS, GNS–PEG, and GNS–PEG–OAm applications. CLSM is a classic yet valuable technique that collects morphological information within cells and tissues. A thorough discussion of the application of this technique on evaluating nanoparticle skin permeation has been previously discussed.³⁹ The FITC-labeled nanoparticles exhibited excitation and emission wavelengths of 492 nm and 510 nm, respectively, as determined by fluorescent spectroscopy.

Cross-section images perpendicular to the skin facilitate the evaluation of GNS permeability into rat skin by observing the distribution pattern of labeled GNS in deep regions of excised skin, including the epidermis, dermis,

and subcutaneous adipose tissue. As shown in Figure 7A, which depicts the CLSM and TEM micrographs of untreated skin, respective layers of the skin can be clearly visualized by both imaging methods. Diffusion profiles into rat skin, as shown in Figure 7B, represent results 24 hours after topical application of the different nanoparticles. FITC-labeled pristine GNS predominantly accumulated on the skin surface near follicle openings, and relatively low fluorescent intensity was observed at the deeper epidermis and dermis layer. Furthermore, strong intensity near the surface of the epidermis suggests that most of the pristine GNS were unable to permeate deeper into the dermis. On the other hand, green fluorescence in the subcutaneous adipose tissues of skins treated with GNS–PEG and GNS–PEG–OAm nanoparticles suggests deeper permeation of functionalized GNS. Localization of GNS–PEG and GNS–PEG–OAm in subcutaneous adipose tissues was further supported by TEM, which has been previously used to access the extent of nanoparticle deposition in skin layers.^{40,41} TEM also revealed that the permeated nanoparticles were well dispersed inside the subcutaneous adipose tissue and eliminated the concerns of fluorescence quenching.

Fluorescence intensity inside different skin layers was quantified and averaged as shown in Figure 8. Both PEG and PEG–OAm functionalization promoted nanoparticle deposition in all skin layers assessed. The highest accumulation of GNS was observed in the epidermis in all groups, indicating that most nanoparticles were limited to the surface layer of the skin due to barrier function of the SC. However, when comparing the fluorescent intensity in the epidermis, one- and twofold increases in deposition were observed upon PEG

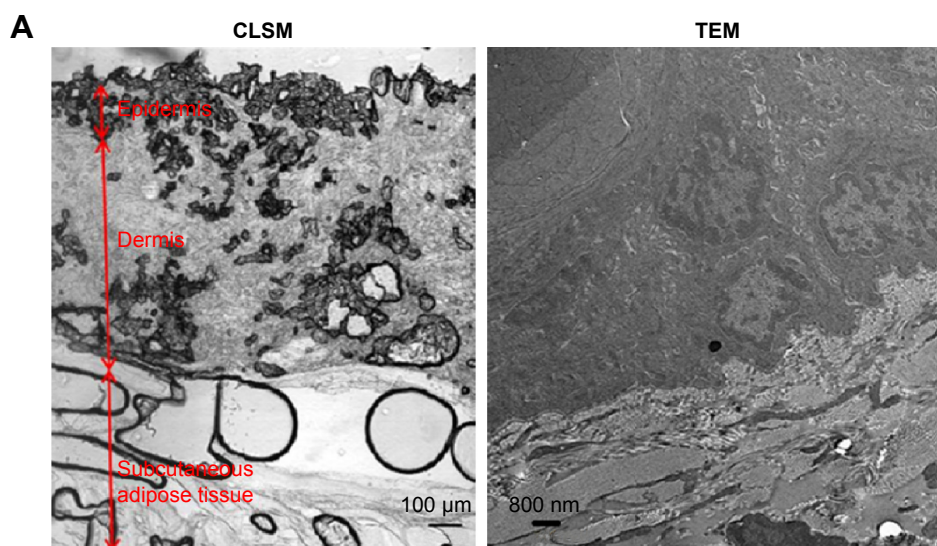


Figure 7 (Continued)

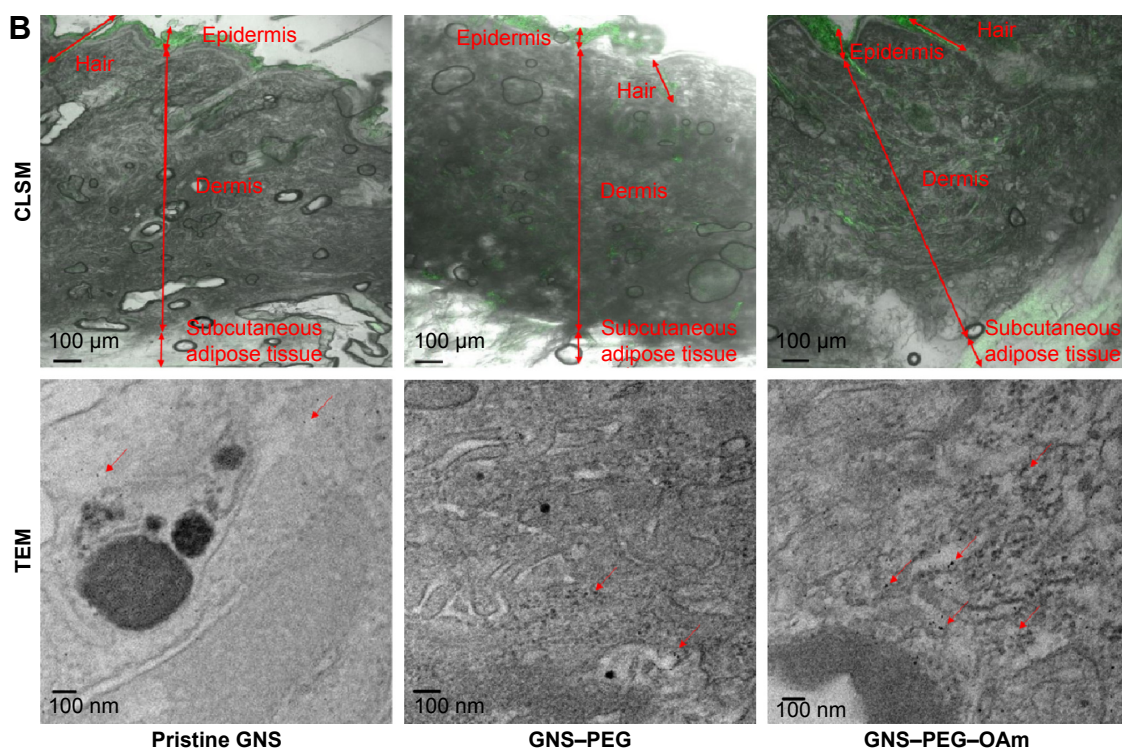


Figure 7 The penetration ability of gold nanoparticle in the skin tissue observed by CLSM and TEM.

Notes: (A) CLSM and TEM images of untreated skin depicting the multiple layer structure. (B) CLSM images of different skin layers and TEM image of the subcutaneous adipose tissue 24 hours after topical application of FITC-labeled pristine GNS, GNS-PEG, and GNS-PEG-OAm. The red arrows indicate the presence of gold nanoparticle. CLSM magnification: 60 \times ; TEM magnification: 5,000 \times .

Abbreviations: CLSM, confocal laser scanning microscopy; FITC, fluorescein isothiocyanate; GNS, gold nanoparticles; OAm, oleylamine; PEG, poly(ethylene glycol); TEM, transmission electron microscopy.

and PEG-OAm functionalization. This reveals that more functionalized GNS were able to cross the SC and become deposited in the deeper epidermis. Skin penetration ability of molecules is based on their chemical structure and interaction with the skin.⁴² PEG has been shown to solvate keratin and extract lipid, thereby disrupting the structural organization and allowing the molecules to pass through.⁴³ Furthermore, PEG

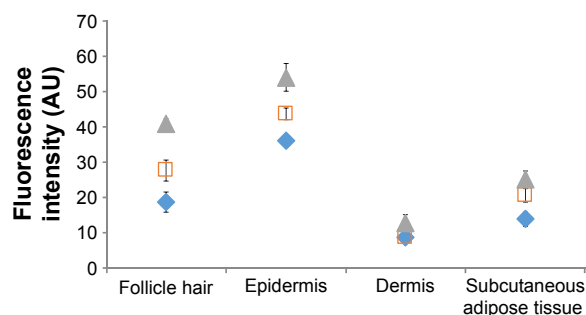


Figure 8 Fluorescent intensity in respective skin layers 24 hours after topical application of FITC-labeled pristine GNS (diamond), GNS-PEG (square), and GNS-PEG-OAm (triangle).

Notes: Fluorescence intensity was calculated by randomly selecting images and extracting z-axis fluorescence intensity from the region of interest using Leica Application Suite X. Error bars represent SD, n=5.

Abbreviations: FITC, fluorescein isothiocyanate; GNS, gold nanoparticles; OAm, oleylamine; PEG, poly(ethylene glycol).

may act as a spacer arm to increase the solubility and allow the terminal OAm to move flexibly in solution, enhancing interaction with the skin.^{44,45} On the other hand, OAm is a cationic lipid and known penetration enhancer that readily interacts with negatively charged skin surface and cell membrane as well as induces conformation changes in surface protein conformation.⁴⁶ In this example, the hydrophilic chain interacts with the polar lipid region, and the hydrophobic chain themselves are inserted into the SC to disrupt the lipid bilayer structure.^{47,48} The functionalized nanoparticles also accumulated in hair follicles in greater amounts, which provided an additional route for transdermal passage. Permeation through the skin compartments in follicles has been shown to be highly in favor of lipophilic vehicles since these compartments are similarly lipid rich.^{49,50} In combination, PEG-OAm functionalization resulted in threefold increase in GNS deposition in the subcutaneous adipose tissue, providing direct proof that PEG-OAm increases permeation of GNS.

Conclusion

This study demonstrates the positive effects of PEG and PEG-OAm functionalization on GNS skin permeation in an

in vivo rat model. Taking advantage of the strong affinity between thiol and gold, PEG and PEG–OAm can be coated onto GNS nanoparticles. Permeation study revealed that PEG- and PEG–OAm-functionalized GNS significantly increased nanoparticle deposition in the deeper layer subcutaneous adipose tissue, which suggests the role of PEG–OAm as a permeation enhancer. On the other hand, pristine GNS was generally incapable of permeating the most superficial layers of the skin. Although the use of human subjects would most accurately reflect true clinical practice, due to regulations, in vivo permeation studies are not always applicable at the initial stages of drug development. An in vivo rat model, in contrast to in vitro models, takes into consideration the structural changes of the skin, blood flow, and metabolism. Thus, this study is meant to provide a preliminary understanding of the effects of selected permeation enhancers on the skin permeation of GNS in vivo. Taken together, this study suggests that PEG and PEG–OAm functionalization is an effective strategy for enhancing the in vivo permeation of GNS for topical administration.

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Disclosure

The authors report no conflicts of interest in this work.

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